

## Chapter 2: Chemistry of Life

### Test Bank

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#### MULTIPLE CHOICE

1. Which subatomic particle has a positive charge?

- A. proton
- B. neutron
- C. electron
- D. nucleus

ANS: A                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

2. Which subatomic particle has no charge?

- A. proton
- B. neutron
- C. electron
- D. nucleus

ANS: B                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

3. Which subatomic particle has a negative charge?

- A. proton
- B. neutron
- C. electron
- D. nucleus

ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

4. Which subatomic particle is found in the nucleus?

- A. proton
- B. neutron
- C. electron
- D. both A and B

ANS: D                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

5. Electrons are found

- A. in the nucleus
- B. in orbitals
- C. at various distances from the nucleus called energy levels
- D. both B and C

ANS: D                      PTS: 1                      DIF: Application      REF: Pages 27-28  
TOP: Atoms

6. The atomic number of an atom is the number of

- A. protons
- B. neutrons
- C. electrons
- D. both A and B

ANS: A                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

7. The atomic mass of an atom is the number of

- A. protons  
B. neutrons
- C. electrons  
D. both A and B

ANS: D                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

8. The subatomic particle that determines how an atom unites with other atoms is the
- A. proton  
B. neutron
- C. electron  
D. both A and B

ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

9. An atom that contains 20 protons, 21 neutrons, and 20 electrons has an atomic number of
- A. 20  
B. 41
- C. 40  
D. 61

ANS: A                      PTS: 1                      DIF: Application      REF: Page 27  
TOP: Atoms

10. An atom that contains 20 protons, 21 neutrons, and 20 electrons has an atomic mass of
- A. 20  
B. 41
- C. 40  
D. 61

ANS: B                      PTS: 1                      DIF: Application      REF: Page 27  
TOP: Atoms

11. An atom that contains 20 protons, 21 neutrons, and 20 electrons has
- A. a positive charge  
B. a negative charge  
C. no charge (electrically neutral)  
D. not enough information is given to determine its charge

ANS: C                      PTS: 1                      DIF: Application      REF: Page 27  
TOP: Atoms

12. Which of these elements is not one of the four elements that make up most of the human body?

- A. carbon  
B. nitrogen
- C. oxygen  
D. calcium

ANS: D                      PTS: 1                      DIF: Memorization  
REF: Page 28              TOP: Elements, molecules, and compounds

13. Bonds that usually dissociate in water to form electrolytes are \_\_\_\_\_ bonds.

- A. ionic  
B. covalent
- C. organic  
D. both B and C

ANS: A                      PTS: 1                      DIF: Memorization  
REF: Page 29              TOP: Ionic bonds

14. The bonds formed when electrons are shared are called

- A. electrolytes
- B. ionic bonds
- C. covalent bonds
- D. inorganic bonds

ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 30              TOP: Covalent bonds

15. The process of dehydration synthesis

- A. uses water to turn large molecules into smaller ones
- B. adds a molecule of water to the reactants
- C. converts smaller molecules into larger ones by removing water
- D. both A and B

ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 31              TOP: Water

16. The process of hydrolysis

- A. uses water to turn large molecules into smaller ones
- B. removes a molecule of water from the reactants
- C. converts smaller molecules into larger molecules by removing water
- D. both B and C

ANS: A                      PTS: 1                      DIF: Memorization  
REF: Page 31              TOP: Water

17. Acids have

- A. a pH less than 7
- B. more  $H^+$  ions than  $OH^-$  ions
- C. more  $OH^-$  than  $H^+$  ions
- D. both A and B

ANS: D                      PTS: 1                      DIF: Memorization  
REF: Page 32              TOP: Acids, bases, and salts

18. Bases have

- A. a pH less than 7
- B. more  $H^+$  ions than  $OH^-$  ions
- C. a pH greater than 7
- D. both A and B

ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 32              TOP: Acids, bases, and salts

19. A solution with a pH of 4

- A. has 100 times more  $H^+$  ions than a solution with a pH of 2
- B. has 100 times fewer  $H^+$  ions than a solution with a pH of 2
- C. has 100 times fewer  $H^+$  ions than a solution with a pH of 6
- D. is basic

ANS: B                      PTS: 1                      DIF: Synthesis              REF: Page 32  
TOP: Acids, bases, and salts

20. The end product of a reaction between a strong acid and a strong base is

- A. water
- B. a salt
- C. a weak acid and a weak base
- D. both A and B

ANS: D                      PTS: 1                      DIF: Memorization

REF: Page 32 TOP: Acids, bases, and salts

21. Which of the following is an example of a monosaccharide?

- A. sucrose
- B. glucose
- C. lactose
- D. glycogen

ANS: B PTS: 1 DIF: Memorization  
REF: Page 33 TOP: Carbohydrates

22. Which of the following is an example of a polysaccharide?

- A. sucrose
- B. glucose
- C. lactose
- D. glycogen

ANS: D PTS: 1 DIF: Memorization  
REF: Page 33 TOP: Carbohydrates

23. Triglycerides

- A. are steroid lipids
- B. have a phosphorus-containing unit on one end
- C. have two fatty acids
- D. have three fatty acids

ANS: D PTS: 1 DIF: Memorization  
REF: Page 34 TOP: Lipids

24. Phospholipids

- A. contain glycerol
- B. contain two fatty acids
- C. contain three fatty acids
- D. are steroid lipids

ANS: B PTS: 1 DIF: Memorization  
REF: Page 34 TOP: Lipids

25. Cholesterol

- A. contains three fatty acids
- B. contains two fatty acids
- C. is a steroid lipid
- D. contains glycerol

ANS: C PTS: 1 DIF: Memorization  
REF: Pages 34-35 TOP: Lipids

26. Which of the following is not true of proteins?

- A. They have water-repelling tails.
- B. They are made up of amino acids.
- C. They contain nitrogen.
- D. They contain peptide bonds.

ANS: A PTS: 1 DIF: Memorization  
REF: Page 34 TOP: Proteins

27. Which of the following is a structural protein?

- A. collagen
- B. keratin
- C. enzymes
- D. both A and B

ANS: D PTS: 1 DIF: Memorization  
REF: Page 35 TOP: Proteins

28. Which of the following is a functional protein?  
A. collagen  
B. keratin  
C. enzymes  
D. both A and B
- ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 36              TOP: Proteins
29. Which of the following substances is not found in a DNA nucleotide?  
A. phosphate unit  
B. glycerol molecule  
C. nitrogen base  
D. a sugar
- ANS: B                      PTS: 1                      DIF: Memorization  
REF: Page 34              TOP: Nucleic acids
30. Which substance is found only in DNA?  
A. adenine  
B. guanine  
C. thymine  
D. cytosine
- ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 36              TOP: Nucleic acids
31. The nitrogen atom has a total of seven electrons. To have a full outer energy level, it would have to  
A. add one electron  
B. lose one electron  
C. add three electrons  
D. lose two electrons
- ANS: C                      PTS: 1                      DIF: Synthesis              REF: Page 27  
TOP: Atoms
32. Which type of chemical bond does not result in the formation of a new molecule?  
A. hydrogen bond  
B. ionic bond  
C. covalent bond  
D. None of the above; all chemical bonds result in the formation of a new molecule.
- ANS: A                      PTS: 1                      DIF: Memorization  
REF: Page 30              TOP: Hydrogen bonds

## TRUE/FALSE

1. Matter is anything that occupies space and has mass.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Levels of chemical organization

2. The mass of an atom is determined by the total number of protons and electrons.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

3. The two subatomic particles found in the nucleus of the atom are protons and neutrons.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

4. A full atomic orbital always contains eight electrons.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 28              TOP: Atoms

5. The atomic number of an atom is the number of protons plus the number of electrons.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

6. The closer an orbital is to the nucleus of an atom, the higher its energy level.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 28              TOP: Atoms

7. An atom with 11 protons, 12 neutrons, and 10 electrons has an atomic number of 11.

ANS: T                      PTS: 1                      DIF: Application      REF: Page 27  
TOP: Atoms

8. An atom with 11 protons, 12 neutrons, and 10 electrons has an atomic mass of 21.

ANS: F                      PTS: 1                      DIF: Application      REF: Page 27  
TOP: Atoms

9. An atom with 11 protons, 12 neutrons, and 10 electrons has a +1 charge.

ANS: T                      PTS: 1                      DIF: Application      REF: Page 27  
TOP: Atoms

10. An element is a substance composed of only one type of atom.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 28              TOP: Elements, molecules, and compounds

11. All molecules are not necessarily compounds.

ANS: T                      PTS: 1                      DIF: Application      REF: Page 28  
TOP: Elements, molecules, and compounds

12. Chemical bonds form when atoms share, donate, or borrow electrons.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 29              TOP: Chemical bonding

13. Ionic bonds result from atoms sharing electrons.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 29           TOP: Ionic bonds

14. When an ionic compound is put into water, it dissociates into ions.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 29           TOP: Ionic bonds

15. Covalent bonds are formed when atoms share electrons.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 30           TOP: Covalent bonds

16. When a covalent compound is put into water, it dissociates into ions.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 30           TOP: Covalent bonds

17. For a compound to be considered an organic compound it must have a C-O or an H-O bond.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 31           TOP: Inorganic chemistry

18. Water is the most abundant organic compound in the body.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 31           TOP: Water

19. The process of dehydration synthesis makes bigger molecules from smaller molecules.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 31           TOP: Water

20. The process of dehydration synthesis has water as one of its end products.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 31           TOP: Water

21. The process of hydrolysis has water as one of its end products.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 31           TOP: Water

22. One of the end products of hydrolysis would have one more hydrogen atom than it did at the beginning of the reaction.

ANS: T                      PTS: 1                      DIF: Synthesis                      REF: Page 31  
TOP: Water

23. Acids have a higher concentration of  $H^+$  ions than  $OH^-$  ions.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 32                      TOP: Acids, bases, and salts

24. Bases have a higher concentration of  $OH^-$  ions than  $H^+$  ions.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 32                      TOP: Acids, bases, and salts

25. A solution with a pH of 8 has more  $H^+$  ions than a solution with a pH of 4.

ANS: F                      PTS: 1                      DIF: Application                      REF: Page 32  
TOP: Acids, bases, and salts

26. A solution with a pH of 5 has more  $H^+$  ions than a solution with a pH of 7.

ANS: T                      PTS: 1                      DIF: Application                      REF: Page 32  
TOP: Acids, bases, and salts

27. A solution with a pH of 2 has 10 times the number of  $H^+$  ions than a solution with a pH of 3.

ANS: T                      PTS: 1                      DIF: Application                      REF: Page 32  
TOP: Acids, bases, and salts

28. When a strong acid and a strong base react, one of the end products is water.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 32                      TOP: Acids, bases, and salts

29. A weak acid almost completely dissociates in water.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 32                      TOP: Acids, bases, and salts

30. When a strong acid and a strong base react, one of the end products is a salt.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 32                      TOP: Acids, bases, and salts

31. A buffer is a substance that resists a sudden change in pH.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 33                      TOP: Acids, bases, and salts

32. The basic unit of a carbohydrate is a monosaccharide.



ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 33           TOP: Carbohydrates

33. A molecule of glucose is larger than a molecule of sucrose.

ANS: F                   PTS: 1                   DIF: Application   REF: Page 33  
TOP: Carbohydrates

34. Sucrose is an example of a disaccharide.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 33           TOP: Carbohydrates

35. Glycogen and starch are both examples of polysaccharides.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 33           TOP: Carbohydrates

36. The process of dehydration synthesis could be used to convert a monosaccharide into a disaccharide.

ANS: T                   PTS: 1                   DIF: Synthesis     REF: Page 31 | Page 33  
TOP: Water and carbohydrates

37. Both fats and oils are lipids.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 34           TOP: Lipids

38. A triglyceride contains two fatty acid molecules.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 34           TOP: Lipids

39. A triglyceride contains a molecule of glycerol.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 34           TOP: Lipids

40. Phospholipids contain three fatty acids.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 34           TOP: Lipids

41. Phospholipids are important molecules in the cell membrane.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Pages 34-35   TOP: Lipids

42. Cholesterol is a steroid lipid.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 35              TOP: Lipids

43. Cholesterol contains two fatty acid molecules.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 34              TOP: Lipids

44. Cholesterol is needed for the formation of several hormones in the body.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 34              TOP: Lipids

45. The basic building block of proteins is nucleotides.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 35              TOP: Proteins

46. The basic building blocks of protein are held together by peptide bonds.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 35              TOP: Proteins

47. Structural proteins include collagen, keratin, and enzymes.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Pages 35-36        TOP: Proteins

48. Enzymes are functional proteins that act as chemical catalysts.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 36              TOP: Proteins

49. The basic building blocks of nucleic acids are nucleotides.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 36              TOP: Nucleic acids

50. The DNA and RNA molecules are the same except the DNA has thymine and the RNA molecule has uracil.

ANS: F                      PTS: 1                      DIF: Application    REF: Page 36  
TOP: Nucleic acids

51. The nitrogen bases adenine, guanine, and cytosine can be found in both RNA and DNA.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 36           TOP: Nucleic acids

52. One difference between DNA and RNA is the type of sugar found in the nucleotides.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 36           TOP: Nucleic acids

53. The smallest unit of matter is the electron.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 27           TOP: Levels of chemical organization

54. The oxygen atom has a total of eight electrons. That means it has six electrons in its outer energy level.

ANS: T                   PTS: 1                   DIF: Analysis           REF: Page 27  
TOP: Atoms

55. The number of electrons in the outer energy level of an atom determines how it behaves chemically.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 27           TOP: Atoms

56. The formula for glucose is  $C_6H_{12}O_6$ . This indicates that there are 24 atoms in a molecule of glucose.

ANS: T                   PTS: 1                   DIF: Application       REF: Page 28  
TOP: Elements, molecules, and compounds

57. The electrolyte most often formed by magnesium (Mg) is  $Mg^{++}$ . This shows that the ion has two more electrons than protons.

ANS: F                   PTS: 1                   DIF: Application       REF: Page 30  
TOP: Ionic bonds

58. Water is the most common solute in the human body.

ANS: F                   PTS: 1                   DIF: Memorization  
REF: Page 31           TOP: Water

59. Both sucrose and lactose are examples of disaccharides.

ANS: T                   PTS: 1                   DIF: Memorization  
REF: Page 33           TOP: Carbohydrates

60. Fats tend to be solids at room temperature.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 34              TOP: Lipids

61. Both cholesterol and phospholipids are involved in the structure of the cell membrane.

ANS: T                      PTS: 1                      DIF: Memorization  
REF: Page 34              TOP: Lipids

62. The lock-and-key model describes how two strands of DNA are able to join so precisely to form a double helix.

ANS: F                      PTS: 1                      DIF: Memorization  
REF: Page 36              TOP: Proteins

## MATCHING

*Match each part of the atom with its corresponding description.*

A. protons    C. electrons  
B. neutrons    D. both protons and neutrons

1. part of the atom that is found in the nucleus
2. part of the atom that is found in orbitals around the nucleus
3. part of the atom that gives an atom its atomic number
4. part of the atom that when combined with the proton gives the atom its atomic mass

1. ANS: D                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms
2. ANS: C                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms
3. ANS: A                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms
4. ANS: B                      PTS: 1                      DIF: Memorization  
REF: Page 27              TOP: Atoms

*Match each organic compound with its corresponding description.*

A. carbohydrates    E. proteins  
B. triglycerides    F. RNA  
C. phospholipids    G. DNA  
D. cholesterol

5. compound whose basic unit is a monosaccharide
6. nucleic acid that contains the nitrogen base uracil
7. lipid that is used to make hormones such as estrogen and testosterone
8. nucleic acid that contains the nitrogen base thymine
9. lipid that is composed of a molecule of glycerol and three fatty acids
10. lipid that has two fatty acids and is important in the cell membrane
11. an enzyme

- |     |              |                    |                   |
|-----|--------------|--------------------|-------------------|
| 5.  | ANS: A       | PTS: 1             | DIF: Memorization |
|     | REF: Page 33 | TOP: Carbohydrates |                   |
| 6.  | ANS: F       | PTS: 1             | DIF: Memorization |
|     | REF: Page 36 | TOP: Nucleic acids |                   |
| 7.  | ANS: D       | PTS: 1             | DIF: Memorization |
|     | REF: Page 34 | TOP: Lipids        |                   |
| 8.  | ANS: G       | PTS: 1             | DIF: Memorization |
|     | REF: Page 36 | TOP: Nucleic acids |                   |
| 9.  | ANS: B       | PTS: 1             | DIF: Memorization |
|     | REF: Page 34 | TOP: Lipids        |                   |
| 10. | ANS: C       | PTS: 1             | DIF: Memorization |
|     | REF: Page 34 | TOP: Lipids        |                   |
| 11. | ANS: E       | PTS: 1             | DIF: Memorization |
|     | REF: Page 35 | TOP: Proteins      |                   |

*Match each term with its corresponding description or definition.*

- |                  |                          |
|------------------|--------------------------|
| A. nucleus       | G. covalent bonds        |
| B. ionic bond    | H. orbitals              |
| C. atomic mass   | I. hydrolysis            |
| D. compound      | J. dehydration synthesis |
| E. electrolyte   | K. acid                  |
| F. atomic number | L. base                  |

12. part of the atom in which electrons are found
13. equal to the number of protons an atom has
14. molecules that form ions when dissolved in water
15. process by which reactants combine only after hydrogen and oxygen atoms have been removed
16. compound that produces  $H^+$  ions
17. part of the atom in which protons are found
18. bond formed by the attraction of atoms or molecules that have opposite charges
19. compound that produces  $OH^-$  ions
20. equal to the number of protons and neutrons in an atom
21. process by which water is used to make smaller molecules form larger molecules
22. bond that is formed when electrons are shared
23. a molecule that contains more than one type of atom

- |     |              |                              |                   |
|-----|--------------|------------------------------|-------------------|
| 12. | ANS: H       | PTS: 1                       | DIF: Memorization |
|     | REF: Page 27 | TOP: Atoms                   |                   |
| 13. | ANS: F       | PTS: 1                       | DIF: Memorization |
|     | REF: Page 27 | TOP: Atoms                   |                   |
| 14. | ANS: E       | PTS: 1                       | DIF: Memorization |
|     | REF: Page 30 | TOP: Ionic bonds             |                   |
| 15. | ANS: J       | PTS: 1                       | DIF: Memorization |
|     | REF: Page 31 | TOP: Water                   |                   |
| 16. | ANS: K       | PTS: 1                       | DIF: Memorization |
|     | REF: Page 32 | TOP: Acids, bases, and salts |                   |
| 17. | ANS: A       | PTS: 1                       | DIF: Memorization |

- REF: Page 27 TOP: Atoms  
18. ANS: B PTS: 1 DIF: Memorization  
REF: Page 30 TOP: Ionic bonds  
19. ANS: L PTS: 1 DIF: Memorization  
REF: Page 32 TOP: Acids, bases, and salts  
20. ANS: C PTS: 0 DIF: Memorization  
REF: Page 27 TOP: Atoms  
21. ANS: I PTS: 0 DIF: Memorization  
REF: Page 31 TOP: Water  
22. ANS: G PTS: 0 DIF: Memorization  
REF: Page 30 TOP: Covalent bonds  
23. ANS: D PTS: 0 DIF: Memorization  
REF: Page 28 TOP: Elements, molecules, and compounds

### SHORT ANSWER

1. Name the three parts of the atom and give a description of each.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 27

TOP: Atoms

2. Explain how an ionic bond forms.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 30

TOP: Ionic bonds

3. Explain how a covalent bond forms.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 30

TOP: Covalent bonds

4. Explain the processes of dehydration synthesis and hydrolysis.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 31

TOP: Water

5. Describe the difference between an acid solution and a base solution in terms of the amount and types of ions in each.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 32

TOP: Acids, bases, and salts

6. Explain the relationship among  $H^+$  ion concentration,  $OH^-$  ion concentration, and pH.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 32

TOP: Acids, bases, and salts

7. Describe the structure of carbohydrates and explain their use in the body.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 33

TOP: Carbohydrates

8. Describe the three types of lipids and give the function of each.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 34

TOP: Lipids

9. Describe the structure of a protein and give examples of a structural protein and a functional protein.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

REF: Page 35

TOP: Proteins

10. Explain the structure of a nucleic acid and list the differences between RNA and DNA.

ANS:

Answers will vary.

PTS: 1

DIF: Memorization

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TOP: Nucleic acids